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appropriate solutions, and gain confidence through practice solving physical chemistry problems.

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Correct answer: I and III.

Explanation: Condition I is always true. Condition II is never true, as Gibbs free energy cannot be negative if enthalpy is positive and entropy is negative.

Condition III may be true if temperature is very high (this is the scenario when the term dominates the term.

Thermochemistry and Thermodynamics - Physical

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## Chemistry Thermodynamics Problems

Problem : Given that the free energy of formation of liquid water is  $-237 \text{ kJ / mol}$ , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate  $\Delta G$  for the reaction, which is  $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$ . Knowing that  $\Delta G = -nFE^\circ$  and  $n = 4$ , we calculate the potential is  $-1.23 \text{ V}$ .

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Entropy, or the amount of disorder, is always highest for gases and lowest for solids. This is because gas molecules are widely spread out and, therefore, are more disordered than solids and liquids. Hydrogen gas will have a higher entropy than liquid water.

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