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cathode (negative electrode) of an electrolytic cell; the electrolyte is a solution of a salt of the metal to be coated; and the anode (positive electrode) is usually either a block ...

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We can write this as the sum of two electrochemical equations  $PbSO_4 + 2e^- \rightleftharpoons Pb + SO_4^{2-}$   $Pb \rightleftharpoons Pb^{2+} + 2e^-$   $?? = ?$

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A Comprehensive Reference for Electrochemical Engineering Theory and Application From chemical and electronics manufacturing, to hybrid vehicles, energy storage, and beyond, electrochemical engineering touches many industries—any many lives—every day. As energy conservation becomes of central importance, so too does the science that helps us reduce consumption, reduce waste, and lessen our impact on the planet. Electrochemical Engineering provides a reference for scientists and engineers working with electrochemical processes, and a rigorous, thorough text for graduate students and upper-division undergraduates. Merging theoretical concepts with widespread application, this book is

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Quantitative methods for the analysis and design of electrochemical systems have progressed greatly over the past forty years. Much of this progress is due to the work of Professor John Newman of the University of California-Berkeley. A tutorial symposium was organized to recognize Prof. Newman's contributions on the occasion of his 70th birthday. This issue contains a series of invited lectures covering the basic principles of electrochemical engineering as well as a variety of examples of applications in electrodeposition, fuel cells, batteries, and electrolytic processes.