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Covalent Bonding and Shapes of. Chapter (c)  $C-Cl \hat{C} + \hat{C}-Cl$  Chlorine is more electronegative than carbon Problem 1.6 Draw Lewis structures, showing all valence electrons, for these molecules.

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Covalent Bonding and Shapes of Molecules 1 CHAPTER 1 Solutions to the Problems Problem 1.1 Write and compare the ground-state electron configurations for each pair of elements: (a) Carbon and silicon C (6 electrons)  $1s^2 2s^2 2p^2$  Si (14 electrons)  $1s^2 2s^2 2p^6 3s^2 3p^2$  Both carbon and silicon have four electrons in their outermost (valence) shells.

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William H. Brown is emeritus professor of chemistry at Beloit College, where he was twice named Teacher of the Year. His teaching responsibilities include organic chemistry, advanced organic chemistry, and, more recently, special topics in pharmacology and drug synthesis.

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Manual to accompany the 7th ed. of the textbook: Organic chemistry by L.G. Wade Jr.

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